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(a) Energy is a broad term that quantifies the amount of work that can be performed. In general, units of energy are expressed as the change in position or temperature of a specific mass of an object.

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To say that a solution has a concentration of 5 M means that in 1 L of solution (not solvent) there would be 5 mol of solute: to prepare such a solution one would place 5 mol of NaCl in a 1-L flask, and then add whatever amount of water is necessary so that the

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Total mass before reaction: $18.00\text{ g} + (25.00 \times 1.025\text{ g/mL}) = 43.63\text{ g}$ After reaction following the law of conservation of mass: $43.63\text{ g} = 12\text{ g} + 30.95\text{ g} + \text{mass of H}_2$
 $\text{mass of H}_2 = 0.68\text{ g}$; volume of H₂ = $0.68\text{ g} \times 1\text{ L} \cdot 0.0824\text{ g} = 8.25\text{ L}$ Solutions Manual for Chemistry Principles and Reactions
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